



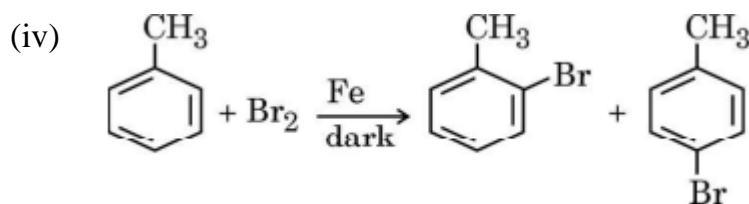
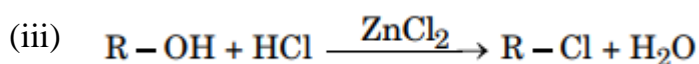
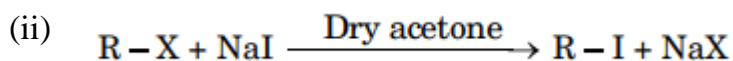
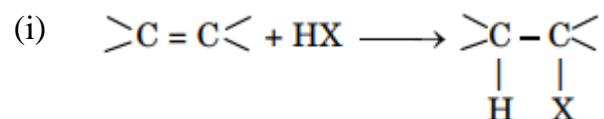
CLASS: XII	DEPARTMENT: SCIENCE 2024-25 SUBJECT: CHEMISTRY	Date of completion: II week of April, 2024
Worksheet No:01 with answers	TOPIC: HALOALKANES AND HALOARENES	NOTE: A4 FILE FORMAT
NAME OF THE STUDENT:	CLASS & SEC:	ROLL NO.

MULTIPLE CHOICE QUESTIONS

1. Which of the following contains sp^2 hybridised carbon bonded to X?

- (i) $CH_2 = CHX$ (ii) $C_6H_5 - CH_2X$
(iii) $CH_2=CHCH_2X$ (iv) $CH_3 - CH_2X$

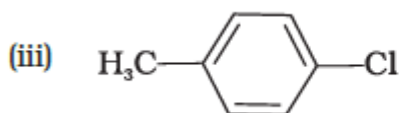
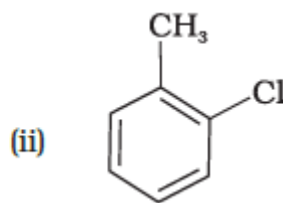
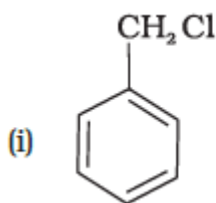
2. Which of the following is a halogen exchange reaction?



3. Toluene reacts with a halogen in the presence of iron (III) chloride giving ortho and para halo compounds. The reaction is

- (i) Electrophilic elimination reaction (iii) Free radical addition reaction
(ii) Electrophilic substitution reaction (iv) Nucleophilic substitution reaction

4. Which of the following isomer has the highest melting point?
- (i) 1,4-Dichlorobenzene (iii) 1,2-Dichlorobenzene
(ii) 1,3 -Dichlorobenzene (iv) All isomers have same melting points
5. Which of the following alkyl halides will undergo S_N1 reaction most readily?
- (i) CH_3CH_2Cl (iii) $(CH_3)_2CHCl$
(ii) CH_3Cl (iv) $(CH_3)_3CCl$
6. Which is the correct IUPAC name for $(CH_3)_3CCH_2Br$?
- (i) 2-Bromo-1,1-dimethylpropane (iii) 1-Bromo-2-methylbutane
(ii) 2-Methyl-1-bromobutane (iv) 1-Bromo-2,2-dimethylpropane
7. The reaction of toluene with chlorine in the presence of iron and in the absence of light yields _____.



(iv) Mixture of (ii) and (iii)

8. Which of the following molecules does not contain a chiral carbon?
- (i) 2-Bromobutane (iii) 2-Bromopropane
(ii) 1-Bromo-1-chlorobutane (iv) 2-Bromopentane
9. The major organic compound formed when 2-Bromobutane is heated with alcoholic KOH is
- (i) Butan-2-ol (iii) 2-Bromopropane
(ii) But-2-ene (iv) But-1-ene
10. Which is the correct increasing order of boiling points of the following compounds?
1-Iodobutane, 1-Bromobutane, 1-Chlorobutane, Butane
- (i) Butane < 1-Chlorobutane < 1-Bromobutane < 1-Iodobutane
(ii) 1-Iodobutane < 1-Bromobutane < 1-Chlorobutane < Butane
(iii) Butane < 1-Iodobutane < 1-Bromobutane < 1-Chlorobutane
(iv) Butane < 1-Chlorobutane < 1-Iodobutane < 1-Bromobutane

Read the given passage and answer the questions that follow:

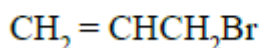
One or more hydrogen atoms of alkanes can be replaced by halogens. Halogenation takes place either at higher temperature (573-773 K) or in the presence of diffused sunlight or ultraviolet light. Free radical chlorination or bromination of alkanes gives a complex mixture of isomeric mono- and polyhaloalkanes, which is difficult to separate as pure compounds. Consequently, the yield of any one compound is low.

11. Among the isomeric alkanes of molecular formula C_5H_{12} , identify the one that on photochemical chlorination yields 4 monochlorides.
12. Is halogenation of alkane in presence of UV an addition or substitution reaction?
13. Identify the final organic product if methane is treated with excess chlorine in UV.
14. How many monochlorides are formed when Butane undergoes halogenation in presence of UV?
15. Name the catalyst used when aryl chlorides are prepared by electrophilic substitution of arenes with chlorine.

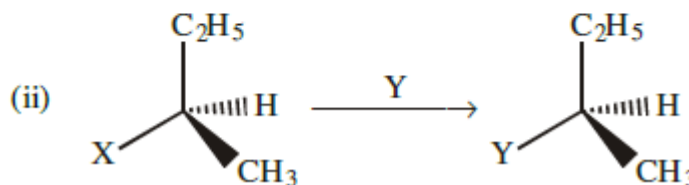
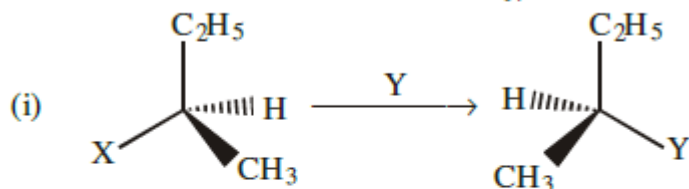
Question – Answer Type:

16. Why is it necessary to avoid even traces of moisture during the use of a Grignard reagent? 1

17. Write the IUPAC name of the following compound: 1

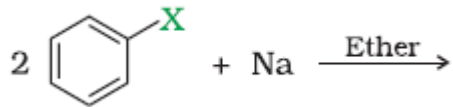
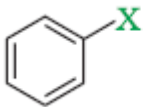
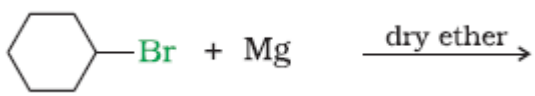
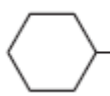


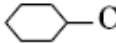

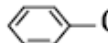
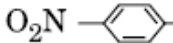
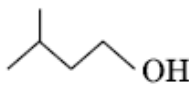
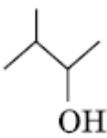
18. Which of the following reactions is S_N1 type? 1



19. Benzyl chloride is highly reactive towards S_N1 reaction. why? 1

20. Arrange the following compounds in order of increasing boiling points. 1
Bromomethane, Bromoform, Chloromethane, Dibromomethane.

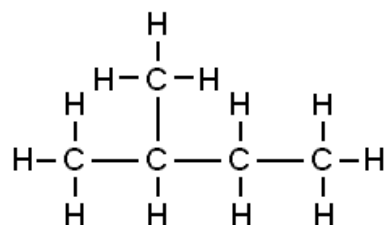
21. Which compound in the following couples will react faster in S_N1 displacement and why? 2
- i) 1-Bromopentane or 2-bromopentane
- ii) 1-Bromo-2-methylbutane or 2-bromo-2-methylbutane.
22. (i) Write the product formed when p-nitrochlorobenzene is heated with aqueous NaOH at 443 K followed by acidification. 2
- (ii) Why dextro and laevo rotatory isomers of Butan-2-ol are difficult to separate by fractional distillation?
23. Out of Chlorobenzene and Cyclohexyl chloride, which one is more reactive towards nucleophilic substitution reaction and why? 2
24. Complete the following reaction: 2
- i) $CH_3Cl + KCN \rightarrow$
- ii) $CH_3OH + SOCl_2 \rightarrow$
25. Give reasons: 2
- (a) Grignard reagent should be prepared under anhydrous conditions
- (b) Alkyl halides are immiscible with water although they are polar.
26. Draw the structures of the major monohalo product for each of the following reactions: 3
- a)  2  + Na $\xrightarrow{\text{Ether}}$
- b)   + Mg $\xrightarrow{\text{dry ether}}$
- c) $H_3C-Br + AgF \longrightarrow$
27. (a) Why are alkyl halides insoluble in water? 3
- (b) Why is Butan-1-ol optically inactive but Butan-2-ol is optically active?
- (c) Although chlorine is an electron withdrawing group, yet it is *ortho*-, *para*-directing in electrophilic aromatic substitution reactions. Why?

28. (a) Out of -Cl and -Cl, which one is more reactive towards S_N2 reaction and why? 3
- (b) Out of -Cl and -Cl, which one is more reactive towards nucleophilic substitution reaction and why?
- (c) Out of  and , which one is optically active and why?
29. Convert the following: 3
- Aniline to Chlorobenzene
 - Bromomethane to Fluoromethane
 - Chlorobenzene to Phenol
30. Among all the isomers of molecular formula C_4H_9Br , identify 3
- the one isomer which is optically active.
 - the one isomer which is highly reactive towards S_N1 .
 - the two isomers which give same product on dehydrohalogenation with alcoholic KOH.
31. Give reasons: 3
- Density of Chloroethane is greater than that of Chloromethane.
 - Boiling points of alkyl halide are higher than their parent hydrocarbon
 - Finkelstein reaction takes place in presence of dry acetone.
32. (a) Why do tertiary alkyl halides undergo S_N1 reaction at a faster rate? 5
- Define Enantiomers.
 - Why is chloroform stored in dark coloured air tight bottles?
 - Nucleophilic substitution of iodoethane is easier than chloroethane. Why?
 - Name the product obtained when 2-chloro-1-phenylpropane is treated with alc. KOH

ANSWERS

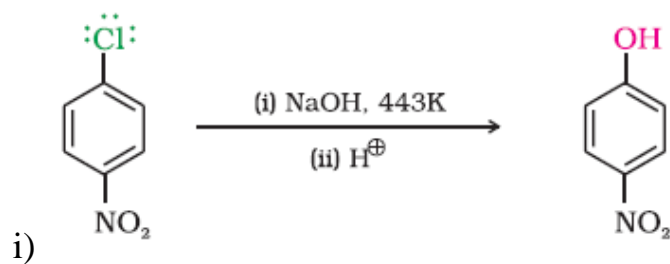
1. (i)
2. (ii)
3. (ii)
4. (i)
5. (iv)
6. (iv)
7. (iv)
8. (iii)
9. (ii)
10. (i)

11.



12. Substitution reaction
13. CCl_4
14. Two.
15. Iron or iron(III) chloride
16. Grignard reagents are highly reactive and react with any source of proton to give hydrocarbons. Even water, alcohols, amines are sufficiently acidic to convert them to corresponding hydrocarbons.
17. 3-Bromoprop-1-ene
18. (ii)
19. Due to the stability of benzyl carbocation as it is stabilized by resonance.
20. Chloromethane < Bromomethane < Dibromomethane < Bromoform
21. i) 2-bromopentane.
It's a secondary haloalkane. Secondary carbocation is more stable than primary.
ii) 2-bromo-2-methylbutane.
It's a tertiary haloalkane. Tertiary carbocation is more stable than primary.

22.



ii) Enantiomers have same boiling points.

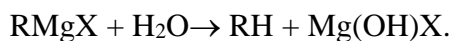
23. Cyclohexyl chloride.

Due to partial double bond character of C-Cl bond in Chlorobenzene / Resonance effect / sp^3 hybridised carbon in cyclohexyl chloride whereas sp^2 carbon in chlorobenzene.

24. i) $CH_3Cl + KCN \rightarrow CH_3CN + KCl$

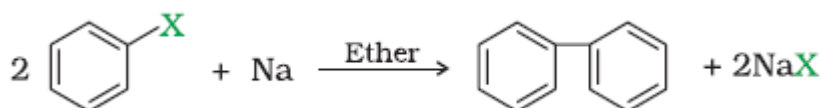
ii) $CH_3OH + SOCl_2 \rightarrow CH_3Cl + SO_2 + HCl$

25. a) It reacts with moisture readily

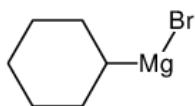


b) They can't form hydrogen bonds with water. Less energy is released when alkyl halide and water are mixed.

26. a)



b)



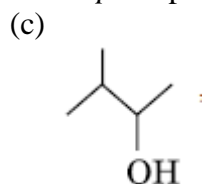
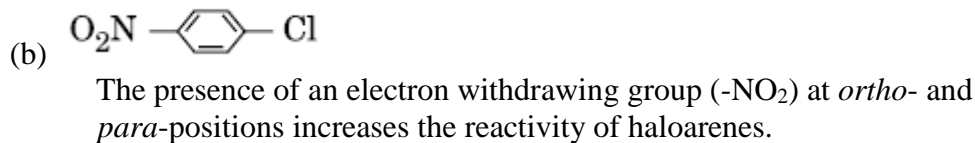
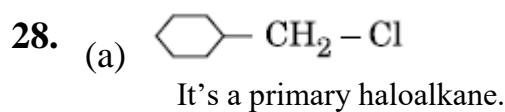
c)



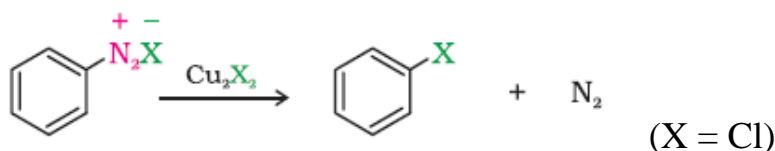
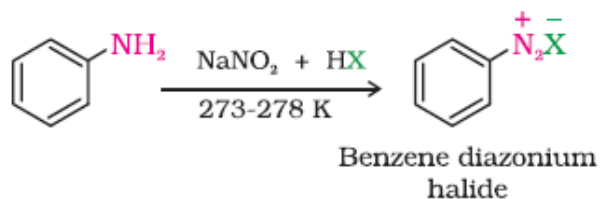
27. a) Haloalkanes are unable to form H-bonds with water molecules. Less energy is released when new attractions are set up between the haloalkane and the water molecules as these are not as strong as the original hydrogen bonds in water.

b) Due to the presence of chiral carbon in butan-2-ol.

c) Due to dominating +R effect over -I effect.



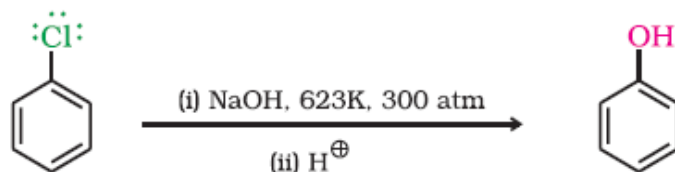
29. i)



ii)



iii)



30. (a) 2-Bromobutane
(b) 2-Bromo-2-methylpropane
(c) 2-Bromo-2-methylpropane and 1-Bromo-2-methylpropane

31. (a) The density increases with increase in number of carbon atoms.
(b) Alkyl halides are polar in nature while hydrocarbons are non-polar / strong dipole-dipole forces in alkyl halides.
(c) NaX formed during the reaction is precipitated in dry acetone and it facilitates the forward reaction.

- 32.** (a) Due to the stability of tertiary carbocation.
- (b) Optically active isomers which are related to each other as non-superimposable mirror images.
- (c) Because chloroform is slowly oxidised by air in the presence of light to an extremely poisonous gas, carbonyl chloride (phosgene).
- (d) Due to the larger size of iodine, bond enthalpy of C-I bond is less than C-Cl bond.
- (e) 1-Phenylpropene

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